Doping a metal (Ag, Al, Mn, Ni and Zn) on TiO$_2$ nanotubes and its effect on Rhodamine B photocatalytic oxidation

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**ABSTRACT**

The effects of ion-doping on TiO$_2$ nanotubes were investigated to obtain the optimal catalyst for the effective decomposition of Rhodamine B (RB) through UV photocatalytic oxidation process. Changing the calcination temperature, which changed the weight fractions of the anatase phase, the average crystallite sizes, the BET surface area, and the energy band gap of the catalyst, affected the photocatalytic activity of the catalyst. The ionic radius, valence state, and configuration of the dopant also affected the photocatalytic activity. The photocatalytic activities of the catalysts on RB removal increased when Ag$^+$, Al$^{3+}$ and Zn$^{2+}$ were doped into the TiO$_2$ nanotubes, whereas such activities decreased as a result of Mn$^{2+}$ or Ni$^{2+}$ doping. In the presence of Zn$^{2+}$-doped TiO$_2$ nanotubes calcined at 550$^\circ$C, the removal efficiency of RB within 50 min was 98.7%.

**Keywords:** Ion-doping, Photocatalytic oxidation, Rhodamine B, TiO$_2$ nanotubes

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1. Introduction

Rhodamine B is a representative of xanthenes cationic dyes and is known for its stability [1]. It is harmful if swallowed and capable to cause irritation to the skin, eyes and respiratory tract. Photocatalytic oxidation is highly efficient in generating ·OH for the degradation of the pollutants [2].

Several techniques have been developed to treat RB. Ozone-ultra-violet (O$_3$/UV) photocatalytic oxidation is highly efficient in generating ·OH for the degradation of pollutants [2]. The choice of photocatalysis is essential in RB removal, and TiO$_2$ is a good candidate. TiO$_2$-photocatalyzed reactions are non-selective oxidation processes in a level of microgram per liter, and the use of TiO$_2$ is effective for inactive substrates [3]. Moreover, the presence of solid particles in TiO$_2$ can provide additional nuclei for bubble formation, thereby enhancing the degradation efficiency of pollutants [4].

Compared with its powder form, TiO$_2$ nanotubes have a larger Brunauer-Emmett-Teller (BET) surface area ($S_{\text{BET}}$), which is advantageous in one-electron oxidation during photocatalytic reactions [5, 6]. However, the relatively large bandgap of TiO$_2$ (3.2 eV) limits the efficiency of photocatalytic reactions because of the high recombination rate of photogenerated electrons and holes [7]. To solve this problem, metal ions were doped into the TiO$_2$ nanotubes to narrow the bandgap of TiO$_2$, as well as reduce the recombination rate of electrons and holes because of the generation of crystal defect caused by ion-doping [7–9]. In addition, the doped ions can act as shallow trapping sites for electrons and holes, and then reacted with H$_2$O$_2$, OH$^-$ and O$_2$ to generate ·OH and O$_2^-$ radicals.

In this study, the effect of metal ion-doping on the photocatalytic activity of TiO$_2$ nanotubes for RB removal via photocatalytic oxidation was investigated. Metal ions, including Ag$^+$, Al$^{3+}$, Mn$^{2+}$, Ni$^{2+}$ and Zn$^{2+}$, were investigated, and Zn$^{2+}$ was observed to be the optimum type. Characteristic analyses of the TiO$_2$ nanotubes were conducted to determine the factors that affect the photocatalytic activities of the catalysts.

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2. Materials and Methods

2.1. Materials

P-25 TiO$_2$ was provided by the Degussa Corporation (Richfield Park, USA). The powder has an $S_{\text{BET}}$ of 50 m$^2$/g, an average crystallite
size of 21 nm, and a crystal structure of 80% anatase and 20% rutile. RB (A.R.) was purchased from Beijing Chemical Reagent Company (China). Microporous membrane (0.45 μm) (MPM) was purchased from Agela Technologies (China).

A photocatalytic reactor (60 mm diameter and 750 mm length), which was equipped with a 37 W 254 nm UV lamp (15 mm Dia. And 793 mm Lth.) purchased from Haili Lighting Equipment Company (China), was used in this experiment. O₂, which was generated by an oxygenator (Yaao, China), was used to produce O₃ with the help of an ozone generator (Tonglin, China). O₃ was guided into the photocatalytic reactor through the bottom port, and the samples were acquired from the sampling port in the middle of the reactor.

2.2. Preparation of TiO₂ Nanotubes

One gram of P-25 TiO₂ was added into 16 mL of a 10 mol/L NaOH aqueous solution in a Teflon vessel. The mixture was then stirred for 2 h at room temperature. The mixture was then heated at 110°C for 24 h in an oven, and cooled to room temperature in air. The as-prepared precipitate (Na₂Ti₃O₇) was washed with distilled water to achieve a pH of 7 and then subsequently dispersed in a 0.1 mol/L HCl aqueous solution. After ultrasonication for 30 min, the precipitate was continuously washed to achieve a pH of 7 and then dried at 60°C. The nanotubes were prepared by calcining the precipitate at 400-600°C for 2 h.

The ion-doped TiO₂ nanotubes were prepared based on the aforementioned process except for adding AgNO₃ (0.0213 g), AlCl₃ (0.0167 g), Mn(NO₃)₂・4H₂O (0.0313 g), NiCl₂ (0.0162 g) or ZnCl₂ (0.0170 g) into the 10 mol/L NaOH solution. The amounts of doping ions which were added in the solution were 1.0% (atom percentage). According to the atlases of energy distribution, the doping concentrations of Ag⁺, Al³⁺, Mn²⁺, Ni²⁺ and Zn²⁺ ions were 0.251%, 0.618%, 0.498%, 0.459% and 0.372%, respectively. A 0.1 mol/L HNO₃ solution was used instead of an HCl solution for the preparation of Ag⁺-doped TiO₂.

2.3. Characterization of TiO₂ Nanotubes

X-ray diffraction (XRD) patterns were collected using a Rigaku Dmax-RB diffractometer (Tokyo, Japan). The surface elemental composition of TiO₂ was recorded via X-ray photoelectron spectroscopy (XPS, AXIS ULTRADLD, Kratos Analytical, UK). Transmission electron microscopy (TEM) images were obtained using a HITACHI HT-7700 electron microscope (Tokyo, Japan). SBET was measured using the Quadrasorb SI-MP apparatus (Quantachrome Instrument, USA). Diffuse reflectance spectroscopy (DRS) analyses were performed using a HITACHI U-3010 UV-vis scanning spectrophotometer (Tokyo, Japan).

2.4. Photocatalytic Activity Tests

Aqueous slurries were prepared by adding 0.15 g of TiO₂-based photocatalyst to 1500 mL of 20 mg/L RB aqueous solution at pH 7. The aqueous slurries were stirred and bubbled with O₂ (1 L/min, 10%) for 30 min and then irradiated with a 254 nm UV lamp. Every 10 min, a 20 mL sample was acquired and mixed with 0.5 mL of 0.5 mol/L Na₂S₂O₃ solution to stop the reaction. The sample was filtered through an MPM to separate the catalyst. The concentration of RB was determined using a UV-visible spectrophotometer (DR5000, Hach).

3. Results and Discussion

3.1. Characterization of Photocatalysts

3.1.1. Morphology of the catalysts

Transmission electron microscopy was performed to observe the morphology of the catalyst. The morphologies of Zn²⁺-doped TiO₂ nanotubes calcined at 550°C are shown in Fig. 1. The TiO₂ nanotube was successfully synthesized. Other ion-doped nanotubes shared a similar morphology (figure not shown). The TiO₂ nanotubes had an average diameter of approximately 10 nm with a hollow and open-ended structure, and the tube wall thickness was approximately 1 nm.

![Fig. 1. TEM images of 550°C calcined Zn²⁺-doped TiO₂ nanotubes.](image)

3.1.2. XRD analysis

The XRD patterns of the un-doped and ion-doped TiO₂ nanotubes are shown in Figs. 3(a) and 3(b), respectively. The weight fractions of anatase in the anatase-rutile mixtures in the samples were calculated using Eq. (1) [10].

\[
f = \frac{1}{1 + 1.265 \frac{I_B}{I_A}}
\]

where f is the weight fractions of anatase in the mixtures, and \(I_B\) and \(I_A\) are the maximum X-ray diffraction peak intensities of rutile phase (110) and anatase phase (101), respectively. The weight fractions of anatase in various samples are listed in Table 1.

Fig. 2 and Table 1 show that both anatase and rutile phases might exist in the TiO₂ nanotubes. When the calcination temper-
nature was increased, the intensity of the anatase phase decreased, whereas that of the rutile phase increased. The dopant ions could substitute Ti^{4+} in the TiO₂ nanotubes lattice, which resulted in the deformation of the crystal lattice. As a result, broadening diffraction peak and decreasing peak intensity were observed.

Among the dopant ions, only Ag⁺ could be observed in the XRD patterns in the form of Ag₂O based on the characteristic peak of the (111) crystal plane of Ag₂O at 32.8° and the XPS spectra of Ag 3 d [11]. Ag⁺ (ionic radius of 115 pm) could be observed because it had difficulty entering the TiO₂ lattice [12]. Ions were partially dispersed in the bulk of TiO₂ nanotubes [9]. The metallic oxides from other doped ions might also be formed on the surface of TiO₂. However, these oxides were not observed in the XRD patterns because their metal sites (0.5 wt.% – 2.0 wt.%) were expected to be below the visibility limit of X-ray analysis [13].

### 3.1.3. XPS analysis

The XPS analysis was performed to determine the surface composition and the chemical state of the elements in the ion-doped TiO₂ nanotubes calcined at 550°C (Fig. 3). Fig. 3(a) shows the XPS survey scan of the un-doped TiO₂, and Figs. 3(b) – 3(k) show the XPS spectra of Ag 3d, Al 2p, Cu 2p, Fe 2p, Mn 2p, Ni 2p, V 2p, Zn 2p, Ti 2p, and O 1s, respectively. The sharp XPS peaks for Ti, O, doped ions, as well as C and Na, are shown in Fig. 3(a). The presence of C was mainly attributed to carbon contamination when the samples were transferred to the XPS device [14]. The presence of Na 1s could be explained by the incomplete exchange of protons with Na⁺ during acid washing [9]. The results also indicate that a part of the catalyst still remained at the Na₂Ti₃O₇ nanotubes where a lot of Na existed [15].

Fig. 3(a) shows that the peak intensities of Ag₂O were much higher than those of other metallic oxides. Moreover, the Ag₂O peak could be observed in the XRD patterns of Ag⁺-doped TiO₂. Both observation results indicate that Ag₂O existed on the surface of the samples. The binding energies of Ag 3d, Al 2p, Mn 2p, Ni 2p and Zn 2p indicate that the doping ions existed as Ag⁺, Al³⁺, Mn²⁺, Ni²⁺ and Zn²⁺ in the TiO₂ crystal lattice (Figs. 3(a) – 3(e)).

Fig. 3(f) shows that binding energies of 458.15 eV to 459.00 eV for TiO₂ nanotubes correspond to the peaks of Ti 2p3/2, whereas binding energies of 464.05 eV to 464.90 eV correspond to Ti 2p1/2. The binding energy of Ti 2p indicates that Ti was incorporated into the lattice of TiO₂ as Ti⁴⁺. The Ti 2p binding energy of the ion-doped TiO₂ nanotubes increased compared with that of pure TiO₂ because the electrons in the conduction band of TiO₂ may be transferred to the doped metal ions on the surface of TiO₂, which resulted in a decrease in the outer electron cloud density of Ti ions [6].

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**Table 1. Catalytic Properties of the TiO₂ Nanotube Samples**

<table>
<thead>
<tr>
<th>Calcination temperature / °C</th>
<th>Weight fractions of anatase phase / %</th>
<th>Indirect energy band gap / (eV)</th>
<th>BET surface areas / (m²/g)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>450 500 550 600</td>
<td>450 500 550 600</td>
<td>450 500 550 600</td>
</tr>
<tr>
<td>un-doped</td>
<td>78.8 82.6 83.4 76.9</td>
<td>3.26 3.24 3.24 3.22</td>
<td>237 169 123 99</td>
</tr>
<tr>
<td>Ag⁺-doped</td>
<td>82.0 83.1 84.6 75.8</td>
<td>3.18 3.13 3.12 3.02</td>
<td>100 62 56 48</td>
</tr>
<tr>
<td>Al³⁺-doped</td>
<td>78.3 86.8 86.0 88.3</td>
<td>3.24 3.23 3.22 3.21</td>
<td>199 143 123 108</td>
</tr>
<tr>
<td>Mn²⁺-doped</td>
<td>92.1 92.2 91.3 90.8</td>
<td>3.05 3.02 3.00 2.90</td>
<td>84 71 61 47</td>
</tr>
<tr>
<td>Ni²⁺-doped</td>
<td>84.0 87.4 88.5 88.8</td>
<td>3.09 3.07 3.06 3.01</td>
<td>89 58 48 46</td>
</tr>
</tbody>
</table>

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**Fig. 2.** XRD patterns of the TiO₂ nanotubes. (a) Un-doped TiO₂ calcined at 400°C to 600°C. (b) Ion-doped TiO₂ with different doped ions calcined at 550°C. ◆ refers to the anatase phase; ◇ refers to the rutile phase, and the words in red indicate the Bragg angles (2θ) of the characteristic peak of (101) crystal planes.
3.1.4. BET surface area analysis

The catalytic efficiencies of the TiO\textsubscript{2} nanotubes are related to the SBET of the catalyst[16, 17]. As summarized in Table 1, the SBET of TiO\textsubscript{2} nanotubes decreased with increasing calcination temperature. The decrease in SBET could be attributed to the aggregation of nanotubes, which resulted in a closely coagulated structure[17].

Table 1 indicates that the SBET of ion-doped TiO\textsubscript{2} nanotubes was more or less smaller than that of un-doped TiO\textsubscript{2} nanotubes, particularly for Ag\textsuperscript{+}, Mn\textsuperscript{2+}, and Ni\textsuperscript{2+}-doped TiO\textsubscript{2} nanotubes. The SBET decreased when metal ions were doped because of partial pore blockages and framework defects [9].

The SBET of TiO\textsubscript{2} nanotubes was significantly larger than that of P-25 TiO\textsubscript{2}. This finding is attributed to the inner and outer
surfaces of the layered-tubular structure, which is one of the special qualities of this catalyst [6].

3.1.5. UV-vis DRS analysis
The UV-vis DRS analysis was performed on the catalysis to obtain the energy band gaps (\(E_g\)) (Fig. 4). The \(E_g\) of the catalysts are calculated [18] and shown in Table 1.

Fig. 4(a) shows that \(E_g\) decreased with increasing calcination temperature. This trend occurred because the \(E_g\) of rutile TiO\(_2\) (3.0 eV) was narrower than that of anatase TiO\(_2\) (3.2 eV), and the rutile phase of TiO\(_2\) began to form when the calcination temperature was increased. The \(E_g\) of un-doped TiO\(_2\) nanotubes was between 3.30 and 3.22 eV, which was a little higher than that of anatase TiO\(_2\) in powder form. This blue shift was due to the quantum size effect on different morphologies of the catalysts [19]. The reflectance spectra of ion-doped TiO\(_2\) slightly shifted toward a longer wavelength (red shift) compared with those of un-doped TiO\(_2\), and the \(E_g\) of ion-doped TiO\(_2\) was visibly narrower. The results show that the doped elements involved were indeed integrated into the lattice of the TiO\(_2\) nanotubes, thereby altering the crystal and electronic structures of the TiO\(_2\) nanotubes [20]. The reduction in \(E_g\) of the TiO\(_2\) nanotubes allows the excitation of the catalyst at a lower irradiation power. Thus, the photocatalytic activity of the catalyst is enhanced [19].

3.2. Photocatalytic Activity of the Catalysts
As shown in Fig. 5(a), the highest catalytic activity for the un-doped TiO\(_2\) nanotubes was observed at a calcination temperature of 500°C, with an RB removal efficiency of 98.1%. The \(S_{BET}\), crystalline phases, crystallinity, and \(E_g\) of TiO\(_2\) nanotubes were considered as crucial factors that affect the photocatalytic activity of the catalyst. When the calcination temperature was increased, the \(S_{BET}\) decreased (see Table 1). The \(S_{BET}\) could affect the photocatalytic activities of the catalyst because a larger \(S_{BET}\) indicates that more pollutants are adsorbed onto the surface of the catalyst, and more areas become available for electron-hole pair separation. The anatase phase exhibited a higher photocatalytic activity compared with other phases. Although the \(S_{BET}\) of TiO\(_2\) nanotubes calcined at 450°C were larger, the crystalline phases were predominantly amorphous, which resulted in low photocatalytic activity, and broadened \(E_g\) of TiO\(_2\). For TiO\(_2\) nanotubes calcined at 550°C, and 600°C, although the main crystalline phases were anatase, the amount of rutile phase increased, and the \(S_{BET}\) decreased; thus, the overall efficiencies were low.

Photocatalytic activity was also affected by ionic radius, valence state, and configuration. The photocatalytic activities of the catalysts on RB removal increased when Ag\(^+\), Al\(^{3+}\) and Zn\(^{2+}\) were doped into the TiO\(_2\) nanotubes, whereas such activities decreased as a result of Mn\(^{2+}\) or Ni\(^{2+}\) doping. In the presence of Zn\(^{2+}\)-doped TiO\(_2\) nanotubes calcined at 550°C, the removal efficiency of RB within 50 min was 98.7%. The photocatalytic activity was affected by the following factors. First, the existing impurity band can reduce the recombination of photoinduced electron-hole, which can transfer electrons more efficiently to the oxygen adsorbed on the surface of TiO\(_2\) nanotubes [21]. Second, the weight fractions of the anatase phases of the catalyst were larger than those of un-doped TiO\(_2\) nanotubes, the \(E_g\) of the nanotubes were narrower, and the \(S_{BET}\) of the catalyst were smaller (see Table 1). Third, the ionic radii of the doped ions affected the photocatalytic activities.

For Ag\(^+\), the ionic radius was significantly larger than that of Ti\(^{4+}\) (60.5 pm). Ag\(^+\) was observed on the surface of TiO\(_2\). The electrons continuously transferred from TiO\(_2\) to AgO, which allowed the pollutant with negative charges to be easily adsorbed by the TiO\(_2\) crystals. The weight fraction of the anatase phase of the catalyst was larger than that of the un-doped TiO\(_2\) nanotubes, and the \(E_g\) of the catalyst was narrower. Therefore, the catalytic activity of Ag\(^+\)-doped TiO\(_2\) increased.

The photocatalytic activity increased when Al\(^{3+}\) was doped into the catalyst. The ionic radii of Al\(^{3+}\) (53.5 pm) was smaller than that of Ti\(^{4+}\); thus, the ions were easily substituted in the lattice. Two doped ions replaced two Ti\(^{4+}\) ions, and one O\(^2-\) hole appeared. The shape of the crystal changed, which facilitated the generation of electron-hole pairs. The weight fraction of the anatase phase of the catalyst increased after ion-doping, and the \(E_g\) was narrower. Therefore, the photocatalytic activities of Al\(^{3+}\)-doped TiO\(_2\) nanotubes increased.
When Mn$^{2+}$ and Ni$^{2+}$ ions were added, the photocatalytic activity of the catalyst decreased. Mn$^{2+}$ and Ni$^{2+}$ ions, whose ionic radii were 67 and 69 pm, respectively, had a slight difficulty in entering the lattice, therefore, MnO and NiO were generated on TiO$_2$ nanotubes. The formation of MnO or NiO could hinder the separation of electrons and holes. In addition, the $S_{\text{BET}}$ of Mn$^{2+}$- and Ni$^{2+}$-doped TiO$_2$ were extremely small. Therefore, the photocatalytic activities of the Mn$^{2+}$- and Ni$^{2+}$-doped catalysts were limited.

The ionic radii of Zn$^{2+}$, which was 74 pm, was slightly larger than that of Ti$^{4+}$. ZnO might be formed on the surface of the TiO$_2$. The combination of TiO$_2$ with the doped metallic oxide results in the generation of Ti$^{3+}$. The electrons were transferred to Ti$^{3+}$ and then to the oxygen, which could reduce the probability for recombination between electrons and holes. In addition, the photogenerated holes left in the valence band would have more opportunity to participate in the oxidizing reactions because photogenerated electrons can be effectively scavenged by Zn$^{2+}$ ions [22]. Therefore, the removal efficiencies increased.

4. Conclusions

Ion-doped TiO$_2$ nanotubes were successfully synthesized for RB removal. The ions were effectively incorporated into the TiO$_2$ lattice, except for Ag$^{+}$. When the ions were doped and the calcination temperature was increased, the $S_{\text{BET}}$ of TiO$_2$ decreased, and the $E_g$ of ion-doped TiO$_2$ became visibly narrower. The photocatalytic
activities of catalysts on RB removal could be increased when Ag\(^{+}\), Al\(^{3+}\) and Zn\(^{2+}\) were doped into the TiO\(_2\) nanotubes, and the Zn\(^{2+}\)-doped TiO\(_2\) nanotubes had the highest catalytic activity because of the effects of the weight fractions of the anatase phase, S\(_{BET}\), \(E_g\) and the doped ions. When Mn\(^{2+}\) and Ni\(^{2+}\) were used as dopants, the photocatalytic activity of the catalyst decreased.

References


